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Hypercoordinate silicon compounds are the focus of intense research due to the diversity of their structures, chemical properties, stereodynamic behaviour and practical use in stereoselective synthesis and medical diagnostics. In recent years a large number of new types of pentacoordinate silicon compounds have been synthesized, including complexes with five different atoms in the silicon environment, compounds with SiO\textsubscript{5}, Si\textsubscript{2}N\textsubscript{2}C, Si\textsubscript{2}O\textsubscript{2}C, Si\textsubscript{2}F\textsubscript{2}C\textsubscript{2}, SiN\textsubscript{4}X (X = S, Se, Te) skeletons and others. At the same time, certain classes of organosilanes containing both penta- and tetracoordinate silicon atoms in the same molecule remain virtually unknown. Among these compounds are \textit{N,N}-bis(dimethylhalogenosilylmethyl)amides, where two silicon centres compete for a single carbonyl group. One of the Si atoms in these amides extends its coordination number to five and forms an (O→Si)-chelate ring while another Si atom remains tetracoordinate. Up to date, very few examples of such compounds have been reported, with the structures of only four complexes determined by X-ray method. Since each of the two silicon atoms in dihalides can potentially form a coordination bond with the carbonyl group, these compounds are particularly interesting as models for studying stereodynamic processes in solutions (such as alternating coordination or permutational isomerisation), pathways of S\textsubscript{2}2–Si reactions, relative contributions of the silicon centres to O→Si coordination and the effects of such coordination on the reactivity of Si\textsuperscript{IV}Me\textsubscript{2}Hal and Si\textsuperscript{IV}Me\textsubscript{2}Hal groups within a single molecule.

Earlier we described (O→Si)-monochelate fluoro silanes RSO\textsubscript{2}-Pro-N(Me)CH\textsubscript{2}SiMe\textsubscript{2}F (5), containing an electron-withdrawing organosulfonyl group at the nitrogen atom of the amino acid fragment. In the present work, we report the synthesis, structures and stereodynamic behaviour of dinuclear fluoro silyl derivatives of proline and sarcosine R\textsubscript{1}R\textsubscript{2}NCH(R\textsubscript{3})C(O)N(CH\textsubscript{2}SiMe\textsubscript{2}F\textsubscript{2})\textsubscript{2} (9), synthesised by bis-silylmethylation of N-organosulfonyl-(S)-proline.
and N-mesylsarcosine amides R₂R₃NCH(R₁)C(O)NH₂ (6) via unstable dichlorides R₂R₃NCH(R₁)C(O)N(CH₂SiMe₂Cl)₂ (7) and isolable N-substituted 2,6-disilamorpholines R₂R₃NCH(R₁)C(O)N(CH₂SiMe₂O)₂ (8).

Discussion of the results

Synthesis of disilamorpholines

Disilamorpholines 8 were prepared by the general synthetic approach developed by us for various silacyclanes. The starting compounds, primary amides 6, were silylmethylated by a mixture of chloro(chloromethyl)dimethylsilane and hexamethyldisilazane with subsequent hydrolysis of unstable dichlorides 7 into target 4-acyl-2,6-disilamorpholines 8 (Scheme 1).

Mesyl and tosyl derivatives of (S)-proline, Ms-Pro-N(CH₂SiMe₂)₂O (8a) and Ts-Pro-N(CH₂SiMe₂)₂O (8b), and mesyl derivative of sarcosine, MsN(Me)CH₂C(O)N(CH₂SiMe₂)₂O (8c), were obtained by one-pot syntheses with yields of 75, 78 and 80%, respectively. The composition and structures of compounds 8 were confirmed by the elemental analysis, IR and multinuclear (¹H, ¹³C, ²⁹Si and CP/MAS ²⁹Si) NMR spectroscopy. The structures of compounds 8a and 8c were also determined by X-ray method (see below).

The formation of hydrolytically unstable dichloride 7a was detected by IR spectroscopy. When a mixture of amide 6a with three equivalents of ClCH₂SiMe₂Cl and one equivalent of (Me₃Si)₂NH was refluxed in benzene or toluene, the absorption of the NCO fragment in 6a was gradually replaced by two absorptions (at 1590 and 1505 cm⁻¹) of the same fragment in 7a, which was typical O→Si chelates of pentacoordinate silicon. IR spectra of all 4-acyl-2,6-disilamorpholines 8a-c showed a strong absorption of the NCO fragment at 1630 cm⁻¹.

In the ¹H NMR spectra of chiral proline derivatives 8a,b, the signals of two SiMe₂ groups appear as four singlets. The ²⁹Si NMR spectra of disilamorpholines 8a–c showed a strong absorption of the NCO fragment at 1630 cm⁻¹. The signals of two SiMe₂ groups appear as four singlets. The ²⁹Si NMR spectra of disilamorpholines 8a–c showed a strong absorption of the NCO fragment at 1630 cm⁻¹.

The above data suggest that both silicon atoms in compounds 8a–c are tetracoordinate. Similar to the double set of signals of SiMe₂ groups in ¹H NMR spectra, the presence of two signals at 75 and 78, respectively. The composition and structures of compounds 8a and 8c were confirmed by the elemental analysis, IR and multinuclear (¹H, ¹³C, ²⁹Si and CP/MAS ²⁹Si) NMR spectroscopy. The structures of compounds 8a and 8c were also determined by X-ray method (see below).

Synthesis of difluorides

In contrast to hydrolytically labile Si–Cl bonds in pentacoordinate dichlorides 7a–c, the Si–F bonds in their difluoro analogues 9a–c were expected to be more stable. (O→Si)-Chelate N’,N’-bis(dimethylfluorosilylmethyl)-N-organosulfonyl-(S)-prolinamides (9a,b) and N’,N’-bis(dimethylfluorosilylmethyl)-N-mesylsarcosinamide (9c) were prepared by the reaction of disilamorpholines 8a–c with BF₃·Et₂O in acetonitrile (Scheme 2).

The composition and structure of difluorides 9a–c were determined by the elemental analysis, IR and multinuclear (¹H,
$^{13}$C and $^{29}$Si NMR spectroscopy. The coordination states of both silicon atoms in compound 9a in the solid state was further confirmed by X-ray single-crystal study (see below) and $^{29}$Si CP/MAS NMR.

**Multinuclear NMR spectroscopy**

$^1$H NMR spectra of difluorides 9a–c contain two signals of the SiMe$_2$ groups in the upfield region. These signals can be attributed to specific SiMe$_2$ groups using Bruker 2D pulse sequence ($^1$H–$^{29}$Si)HMBS. For example, the cross-peaks in the 2D spectrum of 9b (Fig. 1) indicate that the upfield signal of SiMe$_2$ protons corresponds to the signal of pentacoordinate $^{29}$Si at -20 ppm while the downfield signal of SiMe$_2$ protons corresponds to the signal of tetracoordinate $^{29}$Si at +30 ppm.$^{28e}$

Direct spin–spin coupling constants $^1J_{SiF}$ in NMR spectra of compounds 9 for tetracoordinate silicon (230–260 Hz) were generally lower than those for pentacoordinate silicon (ca. 280 Hz; see Experimental section). Such difference, observed both in solutions and solid state, reflected the weakening of the Si–F bond at SiV in comparison with SiIV (see X-ray data for 9a).$^{28e}$ For the same reason, the spin–spin coupling constant $^2J_{HF}$ was observed at ambient temperature only for the SiIVMe$_2$ group but not for the SiVMe$_2$ group. Finally, the weakening of the SiV–F bonds affected the $^2J_{CF}$ constants in $^{13}$C NMR spectra: the observed spin–spin coupling frequencies at SiV centres (10–15 Hz) were significantly lower than those at the SiIV centres (ca. 30 Hz).

Intramolecular O$\rightarrow$Si coordination in complexes 9 in solutions was further confirmed by the downfield shift of the C=O...
signal in their $^{13}$C NMR spectra. The characteristic patterns of Si$^\text{IV}$-silyl and NCH$_2$Si-$^\text{IV}$ signals in $^1$H NMR spectra of difluorides 9a,b (two singlets of equal intensity and an AB-system, respectively) indicated the presence of a chiral carbon atom in their molecules.

The $^{29}$Si signals in solid-state NMR spectra of compounds 9 had greater upfield shifts (ca. $-40$ ppm) than the same signals in solutions (ca. $-10$ ppm). Similar effect was observed for monofluorides 5 and was probably caused by effective solvation of pentacoordinate silicon.

Using the difference between the observed chemical shifts of a Si$^\text{V}$ atom and the typical chemical shift of a Si$^\text{IV}$ atom (ca. 30 ppm), the coordination contribution ($\Delta \delta = \delta_{\text{Si}^\text{V}} - \delta_{\text{Si}^\text{IV}}$) in difluorides 9 can be estimated to be approximately 50 ppm. The comparison of this value to coordination contributions in chlorosilanes RSO$_2$-Pro-N(Me)CH$_2$SiMe$_2$Cl (70–75 ppm), silyloxonium halides [R-Pro-N(Me)CH$_2$SiMe$_2$OH$_2$]X ($X = \text{AlkSO$_2$, ArSO$_2$, Ac}$; $X = \text{Cl, Br}$) (70–80 ppm)$^{18}$ and theoretical data for monofluorosilanes 5 (ref. 7b) and MeC(O)N(Me)CH$_2$SiMe$_2$F$^{11}$ (same as above) indicates a relatively weak coordination in difluorides 9.

All difluorides have two signals in their $^{19}$F NMR spectra: one at approximately $-159$ ppm and another at $-119 \div -125$ ppm. According to literature data, these signals belong to Si$^\text{IV}$Me$_2$F and Si$^\text{IV}$Me$_3$F groups, respectively.$^{2a,e}$

Variable-temperature $^1$H, $^{19}$F, and $^{29}$Si NMR studies

The strength of intramolecular coordination in monochelates of pentacoordinate silicon strongly depends on the nature of the substituent X (Scheme 3; see$^{2a,e}$ and references therein).

In the case of compounds with the OSi$_2$X coordination set and X = Hal or OTf, structures A and B are typical for fluorides, C for chlorides, D for bromides, and E for iodides and triflates.

To study the temperature effects on the coordination set structure in difluorides 9a–c, the temperature-dependent $^1$H, $^{19}$F and $^{29}$Si NMR spectra of these compounds in CDCl$_3$ were obtained. The decrease in temperature from +20 to $-60 \degree \text{C}$ led to reversible downfield shifts of $^1$H and $^{19}$F signals (by ca. 0.03 and 3–4 ppm, respectively) of the Si$^\text{IV}$Me$_2$F group. At the same time, the chemical shift of $^{19}$F in the Si$^\text{IV}$Me$_3$F group was not affected by the temperature. Such behaviour of $^1$H and $^{19}$F signals suggests an increased contribution of form B (Scheme 3) at low temperatures.

Similar to monofluorosilanes,$^{7e}$ the increase in temperature to +60 $\degree \text{C}$ caused very small reversible broadening of the SiMe$_2$ and NCH$_2$ signals in $^1$H NMR spectra of difluorides 9a–c. Such broadening was indicative of a permutational isomerisation at the Si$^\text{IV}$ coordination set of these compounds.

The activation parameters of the permutation were calculated by a $^1$DNMR method using a full line-shape analysis of the signals. For all studied compounds, the stereodynamic processes in CDCl$_3$ were characterised by a narrow range of activation energies (~24 kcal mol$^{-1}$ or greater) and high negative values of the entropy of activation (ca. $\sim -20$ cal mol$^{-1}$ K$^{-1}$). These values were very similar to the activation parameters of N-[dimethylfluorosilylmethyl]- and N-[fluoro(methyl)(phenyl)silylmethyl]amides and -lactams,$^{2a,12-15}$ as well as RSO$_2$-Pro-N(Me)CH$_2$SiMe$_2$F (5),$^{6e}$ where R = Me, Ph, 4-MeC$_6$H$_4$, 4-ClC$_6$H$_4$, 4-BrC$_6$H$_4$, or 4-NO$_2$C$_6$H$_4$.

XRD studies

Disilamorpholine 8a (Fig. 2) crystallizes in two polymorph modifications (8a and 8a$'$).

The orthorhombic (P2$_1$2$_1$2$_1$) crystals 8a were obtained from a heptane–benzene mixture with a molar ratio of 3 : 1, whereas
monoclinic ($P2_1$) crystals 8a’ were obtained from ethanol. There are two crystallographically independent molecules in the asymmetric unit of 8a; its volume is 3.84 times larger than that of 8a’, because the cell 8a’ contains a void of about 40 Å$^3$. The structure of the 2,6-disilamorpholine fragment in compounds 8a, 8a’ and 8c (Fig. 3) is analogous to the previously published five structures (CSD refcodes16 QOMTAN, QOMTER, XATQIT, XULNAT, XULNEX).

The mesyl group and 2,6-disilamorpholine fragment have syn-conformation relative to the proline ring: the corresponding torsion angles C8–N2–S1–C12 and C20–N4–S2–C24 in 8a are 95.3(2)$^\circ$ and 88.8(2)$^\circ$ for two crystallographically independent molecules, respectively, and the torsion angle C8–N2–S1–C12 in 8a’ is 94.7(5)$^\circ$.

An asymmetric unit of difluoride 9a contains two crystallographically independent molecules, which differ by mutual orientation of the proline moiety and Me$_2$FSiCH$_2$ group relative to the chelate ring. In the case of syn-conformation, the interatomic distance S1⋯Si2 is 5.387(1) Å, whereas for anti-configuration the distance S2⋯Si4 is 6.340(1) Å. In 9a (Fig. 4), one of the silicon atoms is pentacoordinated, and its coordination polyhedron is a distorted trigonal bipyramid (axial angles O1–Si1–F1 and O4–Si3–F3 are 172.1(1)$^\circ$ and 171.9(1)$^\circ$, the deviations of Si1 and Si3 atoms from the planes of equatorial substituents toward fluorne atoms are 0.167(1) Å and 0.176(1) Å for two crystallographically independent molecules, respectively).

The structure of coordination polyhedra of Si1 and Si2 atoms in 9a are noticeably different from those in the series of (O→Si)-chelate $N^\prime$-(dimethylfluorosilylmethyl)-$N^\prime$-methyl-$N^\prime$-(organosulfonyl)prolinamides,7a complexes 1 (ref. 6d) and 3 (ref. 6e) (selected bond lengths are given in Table 1).

The structure of 3 differs significantly from other difluorides due to the coordination of amide oxygen atom with the difluoro-

Quantum-chemical studies of the permutational isomerization

To test the applicability of the mechanism (Scheme 4) previously suggested for the permutational isomerisation of $N^\prime$-(dimethylfluorosilylmethyl)amides1 to $N,N$-bis-(dimethylfluorosilylmethyl)amides, we carried out quantum chemical studies of molecule 9a.

At higher temperatures, the equilibrium $B \rightleftharpoons A$ (Schemes 3 and 4) shifts towards the tetracoordinate topomer A. The nucleophilic attack at the Si atom by a fluoride anion (F$^-$) produces pentacoordinate difluoride G, which subsequently loses the F$^-$ anion and forms tetracoordinate intermediate H. The rotation around the Si–CH$_2$ bond produces topomer A’ and

![Fig. 3 Molecular structure of 8c with thermal ellipsoids shown at the 50% probability level.](Image)

![Fig. 4 Molecular structure of 9a with thermal ellipsoids shown at the 50% probability level.](Image)

| Table 1 Selected bond lengths for structures 9a, 1, 3 and prolinamide derivatives |
|---------------------------------------------|--------|--------|------------|
| Si$^{IV}$–F | 1.613(1) | 1.603 | 1.608 | — |
| Si$^{IV}$–F | 1.693(1) | 1.668 | 1.620 | 1.651–1.671 |
| Si⋯O | 2.062(1) | 2.187 | 2.918 | 2.131–2.220 |
finally complex B′ with inverted orientation of the methyl groups at silicon.

According to our previous study, the external fluorine anion can attack tetracoordinated silicon, and the dissociation energy of resulting Si–F bond in gas phase is equal to \( \sim 90 \text{ kcal mol}^{-1} \). Solvation of the \( \text{F}^- \) anion leads to significant decrease of the Si–F dissociation energy. It is reasonable to assume that similar processes can occur in solution of 9a in CDCl₃. Due to the presence of two dimethylfluorosilylmethyl and one bulky tosyl groups, the silicon atoms seem to be less accessible for nucleophilic attack as compared to \( N\)-(dimethylfluorosilylmethyl) amides, where only one dimethylfluorosilylmethyl group is present. Hence, the stereodynamic processes in solution of 9a can be more complex as compared to \( N\)-(dimethylfluorosilylmethyl)amides.⁷

An alternative mechanism can involve the carbonyl group migration from one dimethylfluorosilyl to another (similarly to derivatives urea) (Scheme 5).

In any case, the cleavage of the Si–O coordination bond and the certain conformational changes are necessary to the transfer the carbonyl oxygen atom from one dimethylfluorosilylmethyl to another. Thus, the detailed inspection of these processes can be very useful for understanding the nature of permutational isomerisation in the solution of 9a.

Quantum-chemical calculations of 9a were carried out using Gaussian 03W program.¹⁹ Hybrid PBE0 functional and 6-311G(d,p) basis set were utilized for structure optimization, hessian calculations, relaxed potential energy scans and transition state search. To account for the effect of nonspecific solvation, the PCM model was applied (the value of dielectric constant corresponded to chloroform). All calculations was performed with tight optimization criteria (Opt = tight) and precise grid for computation of two-electron integrals (Int(Grid = Ultrafine)). Molecular graphics was drawn with ChemCraft program.¹⁹ General views of calculated structures, atomic coordinates and total energies can be found in ESI.†

Analysis of potential energy surface for 9a in its isolated molecule and CDCl₃ (PCM calculation) has shown that the presence of two conformational isomers correspond to the cyclic structures (where the Si–O coordination bond is present) and two other conformers belong are acyclic (Si–O coordination bond is absent). According to quantum chemical calculations, the influence of dielectric continuum used in PCM model leads to significant changes in molecular structure of 9a. The most noticeable change is the decrease of Si1⋯O1 distance from 2.36–2.37 to 2.25 Å. Cyclic conformers are more favourable as compared to the acyclic conformers. The difference between two isolated most stable cyclic and acyclic structures is 2.51 kcal mol⁻¹. The use of PCM model for the description of solvation increases this difference to 4.37 kcal mol⁻¹, which is in good agreement with our earlier calculations.⁷ All cyclic conformers can be characterized by the same geometry of coordination
polyhedra of silicon atoms, so their $^{19}$F and $^{29}$Si chemical shifts should be very close.

Other differences are related to mutual orientation of N-organosulfonyl and dimethylfluorosilyl groups. In isolated cyclic and acyclic forms of 9a, these fragments are much closer to each other than in the solution. In two conformers (9a-cyclic2 and 9a-acyclic2, Fig. 5S and 6S, see ESI†), the Si2···O2 distances between one of the SiMe$_2$F groups and the oxygen atom of the sulfonyl group are 3.698 and 3.720 Å, respectively. The optimization of these conformers in terms of PCM model (9a-cyclic2-CDCI$_3$ and 9a-acyclic2-CDCI$_3$, Fig. 7S and 8S†) increases the separation of the above fragments (the Si2···O2 distances become 4.533 and 3.962 Å). Conformers 9a-cyclic and 9a-acyclic (Fig. 1S and 2S†) are stabilized by weak C–H···O bonds between sulfonyl and methyl groups, so the Si2···O2 distances are 3.724 and 4.450 Å. Again, the application of PCM model increases Si2···O2 distances to 4.292 and 4.376 Å (9a-cyclic-CDCI$_3$ and 9a-acyclic-CDCI$_3$, Fig. 3S and 4ST). Thus, the effect of nonspecific solvation prevents the formation of Si2···O2 interactions, so the permutational isomerisation involving the sulfonyl group is unlikely to take place.

The information about the barrier of rotation around Si1–C3 and N1–C4 bonds can be useful to understand the mechanism of permutational isomerisation of 9a. These barriers were calculated by the relaxed potential energy surface scan of CNCO and F1Si1C3N1 torsion angles (the plots of the energy vs. scan coordinate are placed in ESI† [Fig. 9S and 10S]). The value of the rotation barrier around the Si1–C3 bond in isolated molecule 9a is approximately 7 kcal mol$^{-1}$ (Fig. 9ST), so the rotation around the Si–C bond is possible despite the presence of an Si1–O1 coordination bond. In solution, the value of this barrier is even lower than that in isolated molecule (~5.1 kcal mol$^{-1}$). It is not surprising that the rotation around the N1–C4 bond is less favourable than the rotation around the Si1–C3 bond. Firstly, the N1–C3 bond is intermediate between ordinary and double (Table 2). Secondly, the rotation around the N1–C3 bond is attributed to the formation and cleavage of Si1–O1 and Si2–O1 coordination bonds. Our calculation gave the values of 26.3 kcal mol$^{-1}$ for isolated molecule 9a and 24.8 kcal mol$^{-1}$ for its solution in chloroform (Fig. 10ST). These values are very close to the permutational barriers measured for 9a–c by $^{19}$F DNMR study. Thus, the internal rotation can be responsible for the permutational isomerisation of 9a. Additional justification for this assumption was obtained by the localization of transition states (Fig. 11S and 12ST). The modes of negative vibrations (~67.0 and ~62.8 cm$^{-1}$ for isolated molecule and CDCl$_3$ solution, respectively) correspond to the rotation around the N1–C4 bond and formation/dissociation of Si–O coordination bonds. The difference between energies of the most favourable cyclic conformers and transition state is 28.4 and 29.3 kcal mol$^{-1}$ for isolated molecules and solution of 9a, respectively. These values are in agreement with the results of DNMR study. At the same time, the $\Delta S$ value calculated as the difference between the transition state and cyclic isomer is ~2 kcal mol$^{-1}$ K$^{-1}$, which is much lower than the experimental value. In our opinion, this difference can be explained by specific solvation (for instance, H-bonds between CDCI$_3$ and carbonyl or sulfonyl groups, which can be responsible for stabilization of particular conformers).

**In silico estimation of possible pharmacological applications**

Possible applications of synthesized complexes were evaluated by the search for similar compounds with known activities and computational prediction of biological activity based on “structure–activity” relationships (SAR) models. Such analysis provides a reasonable basis for planning further experimental studies of biological activity.

In this study, we used the PubChem structural search for identification of equivalent and similar structures (https://pubchem.ncbi.nlm.nih.gov). The similarity was assessed by the Tanimoto equation and the PubChem dictionary-based binary fingerprint analysis (https://pubchem.ncbi.nlm.nih.gov/search/help_search.html). The search results for similar compounds are shown in Table 3.

According to Table 3, the studied complexes have different similar compounds with variable known activities. No two complexes have the same most similar compound, which could indicate their similar biological potentials.

Computational prediction of biological activity for studied complexes was carried out using SAR-based online services. Possible therapeutic effects and mechanisms of action were evaluated by PASS Online23 (http://www.way2drug.com/ PASSOnline) while the LD$_{50}$ values for acute rat toxicity were estimated by GUSAR Online24 (http://www.way2drug.com/gusar/autocxpredict.html). The results of these predictions are summarized in Table 4.

The prediction results suggest that synthesized compounds may possess cardiovascular and CNS properties. Low levels of predicted acute rat toxicity makes them suitable for all routes of administration.

**Conclusions**

New difluorides R$_2$R$_2$NCH(R$_1$)C(O)(N(CH$_2$SiMe$_2$F)$_2$ (9a–c) with one pentacoordinate and one tetracoordinate silicon atoms were synthesized by silylmethylation of amides R$_2$R$_2$NCH(R$_1$)C(O)NH$_2$ subsequent hydrolysis of unstable intermediates R$_2$R$_2$NCH(R$_1$)C(O)N(CH$_2$SiMe$_2$Cl)$_2$ (7a–c) into 4-acyl-2,6-disilamorpholines R$_2$R$_2$NCH(R$_1$)C(O)N(CH$_2$SiMe$_2$O)$_2$ (8a–c) and the reaction of the latter compounds with BF$_3$·Et$_2$O.
According to IR and NMR data, the $O\rightarrow Si$ coordination in solutions of these compounds was weaker than in the solid state due to effective solvation of the $Si-F$ bond. The absence of spin-spin coupling constants $J_{HF}$ of the methyl groups at $Si^{V}$ and their retention at $Si^{IV}$ indicates a significant weakening of the $Si-F$ bond at pentacoordinate silicon, which favours its ionization. Based on in silico analysis, the synthesized compounds show a potential for pharmacological studies.

### Experimental section

IR-spectra of compounds in solution and in the solid state were recorded on a Bruker Tensor-27 spectrometer using KBr cells and an APR element, respectively. $^{1}H$, $^{13}C$ and $^{19}F$ NMR spectra in CDCl$_3$ and DMSO-d$_6$ were recorded on a Bruker Avance II 300 ($^{1}H$, 300 MHz; $^{13}C$, 75.6 MHz; $^{19}F$, 282.2 MHz) and Jeol JNM-EX400 ($^{1}H$, 400 MHz; $^{13}C$, 100.6 MHz; $^{19}F$, 376.3 MHz) instruments using standard pulse sequences. $^{29}Si$ NMR spectra were recorded using the $^{1}H-29Si$ HSQC pulse sequence supplied with the Bruker Avance II 600 instrument. The $^{1}H$, $^{13}C$, $^{29}Si$ chemical shifts were measured using Me$_4$Si as internal reference. The $^{19}F$ chemical shifts were measured using BF$_3$ as external reference. Negative values are to high field. $^{29}Si$ NMR CP/MAS spectra in the solid state were recorded on a Jeol JNM-EX-400 instrument using 5 mm zirconia rotors and a Doty probe.

The temperature calibration of the NMR spectrometers was performed by measuring the differences in chemical shifts between non-equivalent protons in methanol (-90...+30 °C).

#### Table 3: The search results for similar compounds in PubChem

<table>
<thead>
<tr>
<th>ID</th>
<th>Hits (probability)</th>
<th>The most similar compound with data on patents or activity</th>
</tr>
</thead>
<tbody>
<tr>
<td>8a</td>
<td>0 (90%); 161 (80%)</td>
<td>$N-[[35]S]-1-Methyl-2-oxopiperidin-3-yl\cdot N-(2-oxopropyl)methanesulfonamide$ (CID 58869395) Patent description: sulfonylaminovalerolactams and derivatives thereof as factor Xa inhibitors</td>
</tr>
<tr>
<td>8b</td>
<td>216 (90%)</td>
<td>$\text{4,4'-Ethylenebis(1-methyl-2,6-piperazinedione)}$ (CID 97592) Patent description: novel drug targets to overcome de novo drug-resistance in multiple myeloma; method of reducing amyloid-beta peptide levels using a bisdioxopiperazine; abatement process for contaminants; bisdioxopiperazines and their use as protection agents bis-dioxopiperazines and their use as protection agents</td>
</tr>
<tr>
<td>8c</td>
<td>0 (90%); 19 (80%)</td>
<td></td>
</tr>
<tr>
<td>9a</td>
<td>0 (90%); 74 (80%)</td>
<td>$\text{(1) N-[4-Amino-5-oxo-5-pyrrolidin-1-ylpentyl]methanesulfonamide (CID 17960593)}$ Patent description: alpha-amino acid sulphonamide compounds</td>
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<tr>
<td>9b</td>
<td>59 (90%)</td>
<td>$\text{(2) [2S]-1-[2-[Methyl(methylsulfonyl)amino]ethyl]pyrrolidine-2-carboxamide (CID 57572120)}$ Patent description: quinolinone compounds as 5-HT4 receptor agonists</td>
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<tr>
<td>9c</td>
<td>0 (90%); 10 (80%)</td>
<td>$\text{N,\text{N}-Dimethyl-2-[methyl(methylsulfonyl)amino]acetamide (CID 57682568)}$ Patent description: HIV integrase inhibitors</td>
</tr>
</tbody>
</table>

* Hits – number of similar compounds (≥90% or ≥ 80% Tanimoto index); CID – PubChem Compound ID.
Table 4 Prediction of therapeutic effects and mechanisms of action (PASS Online) and LD<sub>50</sub> values of acute rat toxicity (GUSAR Online)<sup>a</sup>

<table>
<thead>
<tr>
<th>ID</th>
<th>Top 5 predicted therapeutic effects with probability &gt; 50%</th>
<th>Top 5 predicted mechanisms of action with probability &gt; 50%</th>
<th>Predicted LD&lt;sub&gt;50&lt;/sub&gt; values in mg kg&lt;sup&gt;-1&lt;/sup&gt;, type of administration, class of toxicity</th>
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</thead>
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<td>8a</td>
<td>Not predicted</td>
<td>Acetylcholine neuromuscular blocking agent</td>
<td>IP – out of AD, 127, IV, 4 class, 904, PO, 4 class, SC – out of AD</td>
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<tr>
<td>8b</td>
<td>Not predicted</td>
<td>Acetylcholine neuromuscular blocking agent</td>
<td>IP – out of AD, 96, IV, 4 class, 1270, PO, 4 class, SC – out of AD</td>
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<tr>
<td>8c</td>
<td>Spasmolytic</td>
<td>Acetylcholine neuromuscular blocking agent, anaphylatoxin receptor antagonist</td>
<td>IP – out of AD, 156, IV, 4 class, 453, PO, 4 class, 250, SC, 4 class</td>
</tr>
<tr>
<td>9a</td>
<td>Antianginal, multiple sclerosis treatment, antiparkinsonian, neurodegenerative diseases treatment</td>
<td>Phosphatidylinositol 3-kinase C2beta inhibitor, RANTES antagonist, insulin growth factor agonist, insulin like growth factor 1 agonist</td>
<td>IP – out of AD, 80, IV, 4 class, PO – out of AD, 487, SC, 4 class</td>
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<tr>
<td>9b</td>
<td>Antianginal, cardiovascular analeptic, multiple sclerosis treatment, cell adhesion molecule inhibitor</td>
<td>Integrin alpha2 antagonist, calmodulin antagonist, nicotinic alpha4beta4 receptor agonist</td>
<td>IP – out of AD, 101, IV, 4 class, PO – out of AD, SC – out of AD</td>
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<tr>
<td>9c</td>
<td>Antianginal</td>
<td>Anaphylatoxin receptor antagonist, phospholipid-translocating ATPase inhibitor, 2-haloacid dehalogenase inhibitor, glycosylphosphatidylinositol phospholipase D inhibitor, NADPH peroxidase inhibitor</td>
<td>IP – out of AD, 130, IV, 4 class, PO – out of AD, 225, SC, 4 class</td>
</tr>
</tbody>
</table>

* IP – intraperitoneal route of administration; IV – intravenous route of administration; PO – oral route of administration; SC – subcutaneous route of administration; out of AD – compound is out of applicability domain of QSAR models.

and ethyleneglycol (+30…+85 °C).<sup>a</sup> The activational parameters of the permutational isomerisation were calculated using DNMR-SIM software<sup>a</sup> and a modified Eyring equation.<sup>b</sup> In each case, at least twelve temperature points were obtained to achieve a correlation coefficient of 0.997–0.999.

Chloro(chloromethyl)dimethylsilane, (S)-proline hydrochloride, sarcosine and all solvents were purchased from Acros and Sigma-Aldrich. Ethyl esters of N-mesyl-(S)-proline and N-tosyl-(S)-proline were synthesised as described earlier.<sup>10</sup>

**Ethyl-N-mesyl-N-methylglycinate**

Thionyl chloride (83.3 g, 0.27 mol) was added dropwise to a solution of N-methylglycine (44.5 g, 0.50 mol) in absolute ethanol (200 mL). The mixture was refluxed for 5 h, then the volatiles were removed in vacuum. The residue was suspended in an ice-cold mixture of water (20 mL) and diethyl ether (100 mL), and a solution of potassium hydroxide (28.0 g, 50 mL), and the combined organic solutions were evaporated in vacuum. The residue was distilled to afford 13.7 g (70%) of ethyl-N-mesyl-N-methylglycinate with b.p. 144–145 °C (9 torr) and m.p. 34–35 °C. IR spectrum (KBr, ν, cm<sup>-1</sup>): 1750 (C=O), 1360 and 1160 (SO<sub>2</sub>).<sup>1</sup>H NMR spectrum (CDCl<sub>3</sub>, δ, ppm (J, Hz)): 1.25 (3H, t, 3<sup>2</sup> J= 7.3, CH2CH3); 2.77 (3H, s, CH3N); 2.87 (3H, s, CH3CH2); 4.05 (2H, s, NCH3).<sup>13</sup>C NMR spectrum (CDCl<sub>3</sub>, δ, ppm): 8.9 (CH2CH3); 35.3 (CH3CH2); 51.4 (NCH3); 173.9 (C=O). Found, %: C 37.08; H 6.65; N 7.11. C<sub>6</sub>H<sub>12</sub>N<sub>2</sub>O<sub>3</sub>S. Calculated, %: C 36.91; H 6.71; N 7.17.

**N-Mesyl-(S)-prolinamide (6a)**

Ethyl ester of N’-mesyl-(S)-proline (6.6 g, 30 mmol) was stirred with 50 mL of a 25% aqueous ammonia solution for 5 days at ambient temperature. The precipitate formed was isolated by filtration, dried in the open air and used without further purification. Yield 5.5 g (96%), m.p. 156–157 °C (from EtOH), [α]<sub>D</sub><sup>20</sup> = −101.3° (c 1.93, H2O). IR spectrum (KBr, ν, cm<sup>-1</sup>): 3449, 3170 (NH2); 1619 (NCO), 1321 and 1140 (SO<sub>2</sub>).<sup>1</sup>H NMR spectrum (CDCl<sub>3</sub>, δ, ppm (J, Hz)): 1.75–2.0 (2H, m, 3H, s, CH2CH3); 2.87 (3H, s, CH3CH2); 3.25–3.47 (2H, m, 5-C<sub>2</sub>H<sub>3</sub>); 3.96–4.09 (1H, m, 2-CH2); 6.1 and 6.7 (2H, two broad s, NH2).<sup>13</sup>C NMR spectrum (DMSO-d<sub>6</sub>, δ, ppm): 22.0 (Me); 26.0 (C-4); 32.3 (C-3); 50.7 (C-5); 63.5 (C-2); 175.8 (C=O). Found, %: C 37.35; H 6.39; N 14.50. C<sub>6</sub>H<sub>12</sub>N<sub>2</sub>O<sub>3</sub>S. Calculated, %: C 37.49; H 6.29; N 14.57.
**N-Tosyl-(S)-prolinamide (6b)**

Prepared similar to 6a. Yield 6.2 g (93%), m.p. 161–162 °C (from EtOH), [α]_D^25 = -134.6 (c 1.06, H_2O). IR spectrum (KBr, cm⁻¹): 1643 (NCO), 1344, 1156 (SO₂). ¹H NMR spectrum (DMSO-d₆, δ ppm (J, Hz)): 1.31–1.83 (4H, m, 3,4-CH₂); 2.43 (3H, s, CH₃); 3.11–3.32 and 3.35–3.55 (2H, m, 5-CH₂); 3.91–4.01 (1H, m, 2-CH₅); 5.95 and 6.71 (2H, two broad s, NH₂); 7.36 (2H, d, J = 8.3, H Ar); 7.72 (2H, d, J = 8.3, H Ar). ¹³C NMR spectrum (DMSO-d₆, δ ppm): 22.0 (Me); 25.7 (C-4); 31.8 (C-3); 50.9 (C-5); 63.7 (C-2); 129.2 (C-3,5 Ar); 131.4 (C-2,6 Ar); 135.4 (C-1 Ar); 145.8 (C-4 Ar); 175.5 (C=O). Found, %: C 37.58, H 6.27, N 14.50. C₆H₁₂N₂O₃S. Calculated, %: C 37.49, H 6.29, N 14.57.

**N-Mesityl-N-methylglycinamide (6c)**

Prepared similar to 6a. Yield 4.1 g (82%), m. p. 170–171 °C (EtOH). IR spectrum (KBr, cm⁻¹): 3315, 3170 (NH₂); 1657 (NCO), 1320 and 1150 (SO₂). ¹H NMR spectrum (DMSO-d₆, δ ppm (J, Hz)): 2.79 (3H, s, CH₃); 2.87 (3H, s, CH₂); 3.58 and 3.71 (2H, two s, NCH₂); 5.5 and 6.1 (2H, two broad s, NH₂). ¹³C NMR spectrum (DMSO-d₆, δ ppm): 34.3 (CH₃N); 37.1 (CH₃S); 50.9 (NCC(O)); 174.9 (C=O). Found, %: C 29.18, H 5.92, N 16.81. C₆H₁₀N₂O₃S. Calculated, %: C 28.91, H 6.06, N 16.86.

2,2,6,6-Tetramethyl-4-[N-mesityl-(S)-prolinyl]-2,6-disilamorpholine (8a)

A mixture of 6a (0.96 g, 5 mmol), hexamethyldisilazane (0.81 g, 5 mmol), chloro(chloromethyl)dimethylsilane (2.15 g, 15 mmol) and toluene (10 mL) was refluxed for 4 h, then allowed to cool down, and the precipitate formed was filtered out. The remaining solution was evaporated in vacuum, the residue was dissolved in chloroform (30 mL) and stirred with a solution of NaHCO₃ (0.84 g, 10 mmol) in water (10 mL) for 2 h. The organic layer was separated, the aqueous layer was extracted with chloroform (20 mL), and the combined organic solutions were evaporated in vacuum. Recrystallisation of the residue from heptane/benzene (3:1) mixture afforded 1.32 g (75%) of compound 8a with m.p. 121–124 °C and [α]_D^20 = -55.0° (c 1.31, CHCl₃). Found, %: C 41.25, H 7.56, N 7.80, S 9.03. C₁₀H₁₂N₂O₃S. Calculated, %: C 41.11, H 7.48, N 7.99, S 9.15. IR spectrum (KBr, cm⁻¹): 1631 s (C=O), 1325 s, 1148 s (SO₂). ¹H NMR spectrum (CDCl₃, δ ppm): 0.18, 0.19, 0.21 and 0.31 (four s, 12H, 2Si(CH₃)₃); 1.88–2.34 (m, 4H, C₄H₂ and C₄H₃ Pro); 2.7 and 3.42 (dd, 2H, NCH₂Si, J = 15.4 Hz); 2.83 (dd, 2H, NCH₂Si, J = 15.3 Hz); 3.01 (s, 3H, SCH₃); 3.45–3.52 and 3.56–3.63 (two m, 2H, C₂H₂ Pro); 4.80–4.87 (m, 1H, C₄H₂ Pro). ¹³C NMR spectrum (CDCl₃, δ ppm): -0.76 ± 0.00 (m, 2SiMe₂); 24.65 (C₂ Pro); 30.79 (C₃ Pro); 39.9 (SC); 38.1 and 40.29 (two s, NCH₂Si); 47.54 (C₃ Pro); 58.92 (C² Pro); 169.64 (C=O). ²⁹Si NMR spectrum (CDCl₃, δ ppm): 8.0, 10.5.

**Table 5** Crystallographic data and refinement parameters for the structures 8a, 8a', 8c and 9a

<table>
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<th>Structure</th>
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<th>8a'</th>
<th>8c</th>
<th>9a</th>
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<td>C₆H₁₂N₂O₃S₂Si₂</td>
<td>C₁₀H₁₂N₂O₃S₂Si₂</td>
<td>C₁₀H₁₂N₂O₃S₂Si₂</td>
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<td>350.59</td>
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<td>90</td>
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<td>10 571</td>
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<td><strong>ρmin/ρmax (e Å⁻³)</strong></td>
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<td>0.873/−0.494</td>
<td>0.565/−0.358</td>
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</table>
C5H2 Pro); 4.75 ppm (2H, C5H2 Pro); 3.99 and 3.04 (dd, 2H, NCH3Si, JHH 15.0 Hz); 2.99 and 3.04 (dd, 2H, NCH3Si, JHH 15.95 Hz); 3.39–3.46 and 3.50–3.75 (two m, 2H, C5H2 Pro); 4.87–4.92 (m, 1H, C4H Pro); 7.28 and 7.79 (2H, Ar, JHH 8 Hz).

13C NMR spectrum (CDC13, δ ppm): −0.57 ± 0.00 (m, 2SiMe3); 21.47 (Me); 24.75 (3C Pro); 30.89 (15C Pro); 37.96 and 40.41 (two s, NCH3Si); 48.09 (5C Pro); 57.60 (C2 Pro); 127.48 (C2 and C6 Ar), 136.50 (C1 Ar), 143.04 (C4 Ar), 169.52 (C=C = O).

2,2,6,6-Tetramethyl-4-(N-mesitylsarcosinyl)-2,6-disilamorpholine (8c)

Prepared similar to 8a from 0.83 g of 6c. Yield 1.3 g (80%) with m.p. 151–153 °C from heptane–benzene (7:1). Found, %: C 37.28, H 0.46 (dd, 6H, SiIV(CH3)2, 3.30 (two m, 2H, NCH2SiIV); 3.44 (two m, 2H, C5H2 Pro); 4.87 (m, 1H, C4H Pro); 7.20. C18H30F2N2O3SSi2. Calculated, %: C 34.66, H 6.98, N 8.08, S 9.25. IR spectrum (KBr, JCF 16.9 Hz); 41.36–41.47 (m, CH2Si); 48.24 (C3 Pro); 55.45 (C2 Pro); 127.31 (C2 and C6 Ar), 129.64 (C3 and C5 Ar), 135.89 (C1 Ar), 143.74 (C4 Ar), 172.39 (C=O). 19F NMR spectrum (CDC13, δ ppm): −159.47; −119.31. 29Si NMR spectrum (CDC13, δ ppm): −19.1 (d, JSiF 236 Hz), 28.1 (d, JSiF 276 Hz). 29Si NMR CP/MAS spectrum (δ ppm): −32.0 (d, JSiF 251 Hz), 30.7 (d, JSiF 292 Hz).

N',N'-Bis(dimethylfluorosilylmethyl)-N-mesitylsarcosinamide (9a)

Boron trifluoride (0.36 g, 2.5 mmol) was added dropwise to a solution of 8a (0.88 g, 2.5 mmol) in acetonitrile (5 mL). The reaction mixture was refluxed for 2 h, then evaporated in vacuum. The remaining oil was re-extracted with benzene (15 mL), the precipitate was filtered out, and the solution was evaporated in vacuum. The residue was recrystallised from heptane to afford 0.77 g (82%) of 9a with m.p. 100–101 °C. Found, %: C 34.61, H 6.88, N 8.92. C10H24F2N2O3Si2. Calculated, %: C 34.66, H 6.98, N 8.08, S 9.25. IR spectrum (KBr, [nC0 159.15; 119.31. 29Si NMR spectrum (CDC13, δ ppm): −19.1 (d, JSiF 236 Hz), 28.1 (d, JSiF 276 Hz). 29Si NMR CP/MAS spectrum (δ ppm): −32.0 (d, JSiF 251 Hz), 30.7 (d, JSiF 292 Hz).}

N',N'-Bis(dimethylfluorosilylmethyl)-N-tosyl-(S)-prolinamide (9b)

Prepared similar to 9a from 1.1 g of 8b. Yield 0.9 g (80%) with m.p. 87–88 °C (from heptane). Found, %: C 48.23, H 6.80, N 6.15, S 7.20. C13H24F2N2O3Si2. Calculated, %: C 48.18, H 6.74, N 6.24, S 7.15. IR spectrum (KBr, [nC0 1602 s, 1515 w (C=O), 1336 s, 1151 s (SO2). 1H NMR spectrum (CDCl3, δ ppm): 0.18 and 0.22 (two s, NCH3Si); 41.4 (SMe); 49.7 (NCC O); 165.58 (C=O). 19F NMR spectrum (CDCl3, δ ppm): −159.47; −119.31. 29Si NMR spectrum (CDC13, δ ppm): −19.1 (d, JSiF 236 Hz), 28.1 (d, JSiF 276 Hz). 29Si NMR CP/MAS spectrum (δ ppm): −32.0 (d, JSiF 251 Hz), 30.7 (d, JSiF 292 Hz).}

N',N'-Bis(dimethylfluorosilylmethyl)-N-mesitylsarcosinamide (9c)

Prepared similar to 9a from 0.80 g of 8c. Yield 0.73 g (85%) with m.p. 135–136 °C (from heptane). Found, %: C 34.61, H 6.88, N 7.86, S 8.92. C10H24F2N2O3Si2. Calculated, %: C 34.66, H 6.98, N 8.08, S 9.25. IR spectrum (KBr, [nC0 159.15; 119.31. 29Si NMR spectrum (CDC13, δ ppm): −19.1 (d, JSiF 236 Hz), 28.1 (d, JSiF 276 Hz). 29Si NMR CP/MAS spectrum (δ ppm): −32.0 (d, JSiF 251 Hz), 30.7 (d, JSiF 292 Hz).}

The structures were solved by the direct method by XS program and refined by full-matrix least-squares technique against F2 in the anisotropic–isotropic approximation using XL program. Atom H20 in s was located from the difference Fourier maps and refined freely. All remaining hydrogen atoms were placed in geometrically calculated positions and refined in rigid body model (Ueq(H) = 1.2Ueq(CH, CH2), Ueq(H) = 1.5Ueq(CH3)). The Flack parameter confirms (S)-configuration of the proline fragment. Preparation of graphic materials was performed using OLEX2 software package. Crystallographic data for the structural analysis of 8a, 8a', 8c and 9a have been deposited with the Cambridge Crystallographic Data Centre (CCDC nos 1059570–1059573).
Acknowledgements

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References


