Laboratory simulation of Martian atmospheric chemistry

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Laboratory simulation of Martian trace gas chemistry

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Introduction

Recently there has been a great deal of interest in modelling the concentration and distribution of gases within the Martian atmosphere due to the observation of trace species such as methane [1,2] that could hold clues to currently unobserved processes such as volca-

nism or perhaps even life.

Mars Global Circulation Models (MGCMs) can predict both the seasonal and spatial variability of trace gases on Mars. Current models are very effective at predicting the variability due to transport (dynamical) pro-

cesses [3] though there are only two MGCMs in the world that can model the effects of the chemical reac-

tions that act on the gases [4,5].

To improve the accuracy of these models more must be learnt about the mechanisms and rates of these re-

actions under Martian environmental conditions.

The Laboratory

Set-up

Figure 1: Shows a photo of the current laboratory set-up.

Figure 2: Is a schematic diagram of the apparatus.

A: Environmental simulation chamber
A stainless steel vacuum chamber is evacuated and then filled with a Mars mix gas (95% CO₂ + 5% N₂) at a pres-

sure of ~6mbar to approximate a "stan-

dard" Mars atmosphere.
B: FTIR spectrometer
The FTIR measures the IR spectrum of the gasses within the chamber. The spectra are analysed to calculate the concentrations of the gasses being pro-

duced and destroyed. Spectra are taken at 1 minute intervals to examine the evolution of the gasses with time and to ascertain if an equilibrium has been reached.

The UK-LMD MGCM

The UK-LMD MGCM is a state of the art MGCM devel-

oped by researchers from the Laboratoire de Météoro-

ogie Dynamique (LMD), France; the Instituto de Astro-

fisica de Andalucia, Spain and Oxford University and The Open University in the UK. The UK-LMD version utilises a spectral dynamical core [6] and a semi-

lagrangian advection scheme [7,8] developed in the UK along with physical sub-models developed by re-

searchers at LMD [9]. These sub-models include radi-

ative transfer, cloud microphysics and a water cycle [3].

The chemical sub-model contains 16 species (CO₂, CO, O³⁻, O²⁻, O₂, O, H, H₂, OH, HO₂, H₂O₂, CH₄, Ar, N₂, H₂O ice and H₂O vapour) undergoing 42 chemical reactions. The concentration of each gas species is calculated every 30 minutes taking into account advection, verti-

cal diffusion, convection and chemistry plus any ef-

fects of the CO₂ and water cycles.

The Problem

The reaction rates and mechanisms used are assumed to be the same as those for reactions occurring in the Earth's atmosphere but this might not be the case.

Important differences between Earth and Mars:

- The average temperature of Mars is much lower and so reaction rates must be extrapolated to these lower temperatures.
- The bulk constituent on Mars is CO₂ experiments using CO₂ rather than N₂ or O₂ have been shown to in-

crease the efficiency of certain three-body reactions by ~2.5 times [10].
- Basalt dust is an important constituent in the Martian atmosphere but not on Earth so it is possible that some reaction mechanisms haven't yet been observed.

To characterise these effects on rates and mechanisms we must measure reactions under Martian conditions.

An example: The effect of heterogeneous reactions on ozone

Heterogeneous chemistry involves reactions between species in differ-

ent phases. In the Earth's atmosphere reactions that occur on the sur-

face of liquid aerosols and solid particles in the air can be very impor-

tant, particularly to reactive species such as ozone.

The figures to the right show zonally-averaged seasonal plots of ozone column density (µm-atm) on Mars. Figure A is a run using purely gas-

phase chemistry in the chemical sub-model while in Figure B a suite of heterogeneous reactions have been added. These simulations were run using the UK-LMD version of the MGCM and agree with the published results for similar conditions from the LMD MGCM [4,11].

The model predicts the qualitative distribution of ozone well. In both Figures the maximum value of ozone column density occurs during polar winter (Winter solstice is at LS≈90° in the South and L S≈270° in the North). The maxima occur when the concentration of water vapour in the atmosphere is lowest. This is because species that form from the photolysis of water (HO₂=H₂O, H₂O₃) can destroy the ozone.

However, the LMD team found that the predicted ozone distribution did not agree quantitatively when compared with observations from the IR spectrometer aboard SPICAM [11]. The model underpredicted the maxi-

mum concentration at aphelion (LS=71°) and in late summer where the ozone did not increase quickly enough to match the observations. This was remedied by the addition of heterogeneous reactions involving HO₂ occurring on the surface of water ice cloud particles as seen in Figure B.

These are unlikely to be the only heterogeneous reactions of interest. Indeed, reactions that occur on the surface of dust particles will likely be important on Mars though laboratory simulations are required to char-

acterise them as they are not currently well understood.

Summary

The study of trace gas chemistry on Mars is interest-

ing as it offers clues into underlying active pro-

cesses such as volcanism. To understand observations of trace gases they must be modelled by MGCMs that are capable of predicting their seasonal and spatial distribution because of both transport and chemical processes.

Laboratory simulations of chemistry under Martian environmental conditions are needed both to im-

prove the accuracy of current MGCM reaction schemes but also to identify possible new reactions of interest such as ones occurring on the surface of basalt dust particles.

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